Binary Ionic Compound Quiz

For the following compounds, give the correct name:

1. $MgCl_2$
2. AlN
3. Na_3P
4. BBr_3
5. In_2S_3

For the following names, give the correct formula

6. Barium sulfide
7. Sodium nitride
8. Aluminum fluoride
9. Magnesium bromide
10. Potassium oxide
11. Potassium sulfide
12. Magnesium iodide
13. Gallium fluoride
14. Calcium nitride
15. Cesium oxide

Compound Naming Practice- Polyatomic Compounds

For the following compounds, give the correct name

1. Li_3PO_4
2. $NaBrO_3$
3. $Ca(OH)_2$
4. $(NH_4)_2S$
5. $(NH_4)_2SO_4$

For the following names, give the correct formula

6. Potassium sulfate
7. Ammonium hydroxide
8. Iron (III) hydroxide
9. Sodium nitrate
10. Sodium nitrite
11. Ammonium citrate
12. Lithium phosphate
13. Beryllium nitrate
14. Tin (II) hydroxide
15. Ammonium sulfide

Compound Naming Practice- Covalent Compounds

For the following compound, give the correct name

1. Carbon monoxide
2. Carbon dioxide
3. Manganese dioxide
4. Diphosphorous pentoxide
5. Nitrogen dioxide

For the following names, give the correct formula

6. PCl_3
7. N_2O_5
8. SbS_5
9. SiO_2
10. SeF_4

Compound Naming Practice

For each of the following sections; if you're given a name, write the correct formula. If you're given a formula, write the correct name.

Binary Ionic Compounds

1. Potassium bromide
2. Calcium oxide
3. Aluminum chloride
4. Sodium sulfide
5. Potassium iodide
6. CaS
7. K_2O
8. SrI_2
9. Li_3N
10. Al_2O_3
11. Copper (II) chloride
12. Copper (III) nitride

Polyatomic Ionic Compounds

13. Sodium phosphate
14. Ammonium iodide
15. Potassium hydroxide
16. Aluminum nitrate
17. Calcium bromate
18. Na_3PO_4
19. $Ca_3(C_6H_5O_7)_2$
20. $(NH_4)_3N$
21. $B(CH_3COO)_3$
22. $Be(OH)_2$

Covalent Compounds

23. NO_2
24. CO
25. CO_2
26. N_2O_2
27. CCl_4

Balancing Equations Review

Balance each of the equations below and ***state the type of reaction shown***.

1. ____ N_2 + ____ H_2 → ____ NH_3
2. ____ $KClO_3$ → ____ KCl + ____ O_2
3. ____ NaCl + ____ F_2 → ____ NaF + ____ Cl_2
4. ____ H_2 + ____ O_2 → ____ H_2O
5. ____ $Pb(OH)_2$ + ____ HCl → ____ H_2O + ____ $PbCl_2$
6. ____ $AlBr_3$ + ____ K_2SO_4 → ____ KBr + ____ $Al_2(SO_4)_3$
7. ____ Ag_2O → ____ Ag + ____ O_2
8. ____ S_8 + ____ O_2 → ____ SO_3
9. ____ HNO_3 + ____ $NaHCO_3$ → ____ $NaNO_3$ + ____ H_2O + ____ CO_2
10. ____ NaBr + ____ CaF_2 → ____ NaF + ____ $CaBr_2$